

## Removal of Cefixime from Water Using Rice Starch by Response Surface Methodology

Fatemeh Sadat Tabatabaei, Mahdi Asadi-Ghalhari<sup>\*</sup>, Rahim Aali, Fatemeh Mohammadi, Roqiye Mostafaloo, Rezvaneh Esmaeili, Zohreh Davarparast, and Zahra Safari

Department of Environmental Health Engineering, Faculty of Health, Qom University of Medical Sciences, Qom, Iran

### Abstract

**Background:** Remaining pharmaceutical compounds cause environmental pollution. Therefore, refining these compounds has become a major challenge. In this study, the function of eliminating Cefixime (CFX) using rice starch was evaluated under controlled conditions.

**Methods:** Response Surface Methodology (RSM) was used to design, analyze, and optimize experiments, and the interaction between four variables including pH (3-9), rice starch dose (0-300 mg/L), CFX initial concentration (0-16 mg/L) and time (20-120 min) was investigated on CFX removal.

**Results:** The optimum pH, starch dose, initial concentration and time were 4.5, 225 mg/L, 7.9 mg/L and 95 min, respectively. The maximum efficiency of CFX removal was 70.22%. According to RSM, this study follows a quadratic model ( $R^2=0.954$ ).

**Conclusion:** Rice starch has been successful in removing CFX from the aqueous solution. Therefore, it is recommended to utilize this process to remove CFX from aqueous solutions.

*Avicenna J Med Biotech 2020; 12(4): 230-235*

**Keywords:** Cefixime, Response surface methodology (RSM), Rice starch

**\* Corresponding author:**  
Mahdi Asadi-Ghalhari, Ph.D.,  
Department of Environmental  
Health Engineering, Faculty of  
Health, Qom University of  
Medical Sciences, Qom, Iran  
Tel: +98 25 37842227  
Fax: +98 25 37842227  
E-mail:  
mehdi.asady@gmail.com  
Received: 9 May 2020  
Accepted: 27 Jul 2020

### Introduction

In recent years, pharmaceutical compounds have been considered as the most serious water pollutants<sup>1</sup>. One of the most important groups of pharmaceuticals are antibiotics that are used to prevent and treat infections of farms, cattles and human<sup>2</sup>. Water pollution is the result of uncontrolled use and a serious environmental threat, which may lead to creation of antibiotic-resistant microbes and pose a serious threat to all humanity in coming years<sup>3</sup>. Also, these compounds are toxic for aquatic organisms even at low concentrations<sup>4</sup>. The excretion of active and unchanged compounds has been proven in many antibiotics<sup>5</sup>. Since removing these compounds entirely is not feasible in conventional wastewater treatment plants, these compounds enter into surface water, groundwater, and water treatment plants ultimately, and as a result of not being removed from treatment plant, they enter drinking water<sup>6</sup>.

One of the most important antibiotics utilized to treat infection is cefixime. Cefixime is a semi-synthetic antibiotic from the third generation of cephalosporins. Cefixime can be consumed against a variety of bacterial organisms and infections including Staphylococci, Influenza, Hemophilia, *Escherichia coli*, Streptococcus, fever and chills, and throat infections<sup>7,8</sup>. Studies show this compound has been one of the most widely consumed drugs in Europe in recent years<sup>9</sup>.

The molecular structure of Cefixime is demonstrated in figure 1<sup>10</sup>.

Due to persistent incapability of conventional wastewater treatment methods to eliminate or diminish the number of undesired antibiotics, several more efficient treatment methods<sup>11</sup> based on advanced oxidation, ion exchange and adsorption with activated carbon, reverse osmosis and biological treatment are applied among which biological treatment is not very effective in removing antibiotics<sup>7</sup>. The adsorption process is extensively used to remove organic contaminants. This method has an advantage over other chemical methods due to the lack of metabolite production. The adsorp-

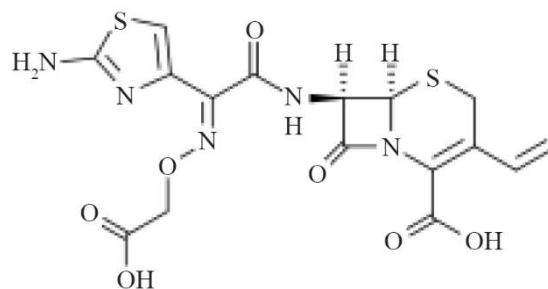


Figure 1. Cefixime molecular structure.

tion process has received much attention compared to other purification techniques regarding initial cost, reuse of effluent, simplicity, and flexibility in design, easy operation and non-sensitivity to pollutants and toxic compounds. Studies have proven that adsorption process is advantageous for removing antibiotics from aqueous environments <sup>12</sup>. In many studies, rice starch has been described as a highly efficient adsorbent to remove heavy metals such as zinc, lead, copper, and cadmium <sup>13</sup>. The average size of rice starch granules is 3-8  $\mu$ l and they are angular and polyhedral. Amylopectin is the dominant component in rice starch and plays a significant role in its physico-chemical properties. Rice starch has more priorities than other starches due to its characteristics such as high stability, high acid resistance and wide range of amylase to amylopectin ratio <sup>14</sup>. Owing to water hygiene protection and human protection against severe health consequences of antibiotics, these compounds must be effectively and appropriately removed from sewage and water resources. The use of natural materials is of great importance in eliminating pollutants such as antibiotic compounds from an environmental point of view.

Studies have shown that the process of removing cefixime from water sources with rice starch has not been investigated so far. In this study, removal of cefixime with rice starch was performed through Response Surface Methodology (RSM). RSM is a statistical method for designing experiments that is used to optimize, evaluate the interaction of independent factors and reduce the number of experiments in chemical and biochemical processes <sup>15</sup>. The objective of using RSM is to identify optimal operating conditions or to enhance operating conditions <sup>16</sup>. In this study, four independent variables including CFX initial concentration, pH, starch dose and time were selected and evaluated at five levels (- $\alpha$ , -1, 0, +1 and + $\alpha$ ). Design Expert software version 12 was used to design the experiment.

### Materials and Methods

#### Chemicals and equipment

All the chemicals used in this study had a laboratory grade. Cefixime (CFX) was prepared with chemical formula of C<sub>16</sub>H<sub>15</sub>N<sub>5</sub>O<sub>7</sub>S<sub>2</sub> and a purity of 0.98% from Sigma-Aldrich. Rice starch, HCL and NaOH were prepared from Merck. 1 g/L CFX stock was made and intended concentrations were obtained from it. In this study, 3% starch solution was prepared using deionized water. To prevent hydration of starch solution during

experiments, the solution was freshly prepared. Stoke solution with suitable dilution ratio was used to make other concentrations used in this study. Evaluating CFX removal function by rice starch was the experiment design which was modeled and optimized using Design Expert software (Version 12). The Central Composite Design (CCD) based on RSM was used to find the optimal conditions, the effect of parameters and their interaction in CFX removal by starch. In this study, CCD with four variables including pH, starch dose, CFX initial concentration and time, at five levels (- $\alpha$ , -1, 0, +1, + $\alpha$ ) was chosen. According to the equation  $2k+2k+C$ , total number of experiments designed was 30 runs which K as the number of factors and C as the number of central points <sup>17</sup>. Tables 1 and 2 show process independent variables, their values and experiment design matrix. The one-way analysis of variance (ANOVA) was utilized to determine adequacy of the proposed model, to determine R<sup>2</sup>, adjusted R<sup>2</sup> and predicted R<sup>2</sup> <sup>18</sup>.

#### Adsorption experiments

To determine relationship between four independent factors of pH (A), starch dose (B), initial CFX concentration (C) and retention time (D) in CFX removal by rice starch, CCD was used. The experiments were performed in 30 ml Erlenmeyer flasks at initial CFX concentrations, and pH, starch dose and retention time were designed by Design Expert. To separate starch from solution, the samples were centrifuged at 4000 rpm for 10 min. The pH of the samples was adjusted using 0.1 N solutions of either HCl or NaOH and measured by pH meter (Inolab pH=7110 model, WT-W). CFX concentration was determined using spectrophotometer (Cecil 7250 Model) at a wavelength of 288.5 nm. Finally, the removal efficiency of CFX was calculated by Eq.1:

$$(1) \quad RE (\%) = [(C_0 - C_t) / C_0] / 100\%$$

Where RE is the removal efficiency of CFX (%), and C<sub>0</sub> and C<sub>t</sub> show initial and final concentrations of CFX (mg/L), respectively.

### Results

#### Statistical analysis

Based on statistical analysis (Table 3), one-way analysis of variance (ANOVA) shows the significance of model in elimination of CFX by different parameters (pH, starch dose, initial CFX concentration and retention time). Also, based on F-value (22.41), the model

Table 1. Levels of independent variables in experiment design

Factor (Unit)	Code	Level				
		- $\alpha$	-1	0	+1	+ $\alpha$
pH	X <sub>1</sub> (A)	3	4.5	6	7.5	9
Starch dose (mg/L)	X <sub>2</sub> (B)	0	75	150	225	300
Initial CFX concentration (mg/L)	X <sub>3</sub> (C)	0	4	8	12	16
Retention time (min)	X <sub>4</sub> (D)	20	45	70	95	120

Table 2. Experimental design matrix used to evaluate the rate of CFX removal by rice starch

Std order	pH	Starch dose (mg/L)	Initial CFX concentration (mg/L)	Retention time (min)	% Removal		
					Obtained	Predicted	Residual
1	4.5	75	4	45	43.00	35.42	83.92
2	7.5	75	4	45	39.48	38.06	2.16
3	4.5	225	4	45	57.49	65.16	42.91
4	7.5	225	4	45	47.40	51.70	-5.67
5	4.5	75	12	45	51.66	60.35	-17.46
6	7.5	75	12	45	70.00	89.89	51.26
7	4.5	225	12	45	64.28	80.62	23.27
8	7.5	225	12	45	73.22	94.07	80.31
9	4.5	75	4	95	35.08	39.26	-74.97
10	7.5	75	4	95	35.96	27.20	58.31
11	4.5	225	4	95	66.33	84.18	30.32
12	7.5	225	4	95	51.49	56.02	22.80
13	4.5	75	12	95	50.48	47.78	87.24
14	7.5	75	12	95	51.80	62.63	-37.58
15	4.5	225	12	95	64.57	83.24	3.17
16	7.5	225	12	95	63.69	81.98	-2.34
17	3	150	8	70	52.67	65.02	-45.74
18	9	150	8	70	53.68	66.40	-41.16
19	6	0	8	70	45.09	50.54	-32.98
20	6	300	8	70	69.64	99.63	-53.93
21	6	150	0	70	0.00	37.03	-36.43
22	6	150	16	70	46.47	54.59	-50.48
23	6	150	8	20	60.28	83.58	-86.89
24	6	150	8	120	60.50	75.33	-0.02
25	6	150	8	70	62.92	76.38	38.04
26	6	150	8	70	61.16	76.38	2.63
27	6	150	8	70	62.26	76.38	24.69
28	6	150	8	70	62.04	76.38	20.26
29	6	150	8	70	59.84	76.38	-23.52
30	6	150	8	70	57.86	76.38	-62.11

was found to be significant which means only 0.01% of data variables are likely not to be considered by the model. The Prob>F value was less than 0.05, which means the model statements are significant. The F-value for Lack of Fit in this study was 4.18, demonstrating that the relationship between Lack of Fit and pure error was not significant. In this study, the terms B, C, CD, BD, AD, AC, AB, C<sup>2</sup> were significant (p<0.05). The results obtained for the analysis of cefixime removal are presented in table 3. The values of regression coefficient (R<sup>2</sup>) and adjusted regression coefficient (Adj. R<sup>2</sup>) for the predicted quadratic model were 0.954 and 0.911, respectively.

The regression coefficient predicted by the model (Pred. R<sup>2</sup>) was also equal to 0.758 which indicates experimental values were consistent with predicted value of the model. The value of Adeq precision in predicted quadratic model was 20.7 which is a value greater than

4, indicating the accuracy of the predicted model. According to table 3, the maximum to minimum response ratio is greater than 3. Hence, to improve the model and consider the most appropriate power transmission function on the response, the Box-Cox test was applied. The Box-Cox diagram brings the data distribution closer to normal distribution. In this test, the best values of Lambda and constant K were chosen to be 1.61 and 0.73, respectively. Therefore, the improved model is demonstrated as a quadratic equation in Eq. 2:

$$(2) \quad (R+0.73)^{1.61} = +763.81 + 3.45 * A + 122.74 * B + 127.23 * C - 20.63 * D - 40.25 * A * B + 67.27 * A * C - 36.74 * A * D - 23.66 * B * C + 37.96 * B * D - 41.01 * C * D - 26.68 * A^2 - 3.23 * B^2 - 118.08 * C^2 + 7.70 * D^2$$

**Effect of process parameters on CFX removal efficiency**

3D contours and perturbation plot were evaluated to assess the effect of variables on CFX removal efficien-

Table 3. One-way analysis of variance (ANOVA) for CFX removal by rice starch

Source	Sum of squares	df	Mean square	F-value	p-value	Prob >F
Model	1347602	14	96257.26	22.41	< 0.0001	Significant
A-pH	286.1759	1	286.1759	0.07	0.7998	
B-starch	361573.9	1	361573.9	84.17	< 0.0001	
C-concentration	388483.1	1	388483.1	90.44	< 0.0001	
D-time	10214.37	1	10214.37	2.38	0.1439	
AB	25923.76	1	25923.76	6.03	0.0267	
AC	72401.92	1	72401.92	16.85	0.0009	
AD	21600.35	1	21600.35	5.03	0.0405	
BC	8960.239	1	8960.239	2.09	0.1692	
BD	23054.48	1	23054.48	5.37	0.0351	
CD	26902.82	1	26902.82	6.26	0.0244	
A^2	19519.67	1	19519.67	4.54	0.0500	
B^2	286.1087	1	286.1087	0.07	0.7999	
C^2	382438.1	1	382438.1	89.03	< 0.0001	
D^2	1625.161	1	1625.161	0.38	0.5477	
Residual	64435.39	15	4295.693			
Lack of fit	57550.4	10	5755.04	4.18	0.0640	Not significant
Pure error	6884.991	5	1376.998			
R-squared				0.9544		
Adj R-squared				0.9118		
Pred R-squared				0.7582		
Adeq precision				20.700		

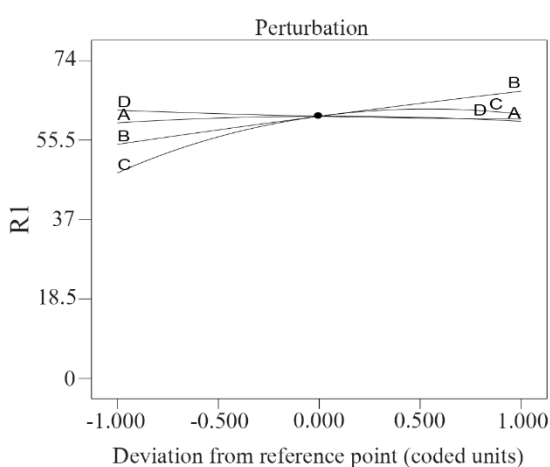


Figure 2. Perturbation plots for the effect of CFX removal. A) PH, B) starch dose, C) initial CFX concentration, D) retention time.

cy using rice starch. As figure 2 shows, the perturbation plot was used to compare the effect of independent factors at a specific point in the design location. In this graph, the slope of each line indicates the effect of that factor on the amount of pollutant removal. In this plot, factors B (Starch dose) and C (Initial CFX concentration) were most effective. According to Eq 2, the initial CFX concentration with a coefficient of 127.23 had the most positive effect on CFX removal.

3D contours in figures 3A and B show the interactions between pH, initial CFX concentration, starch

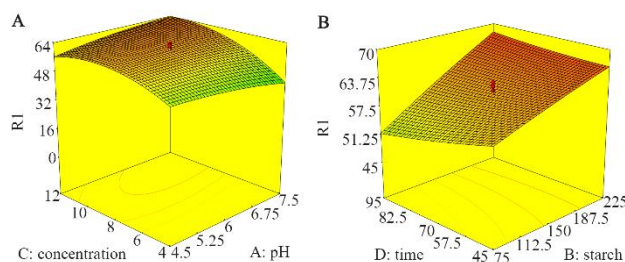


Figure 3. Response surface plots, removing CFX, A) pH and initial CFX concentration, B) starch dose and retention time.

dose and retention time in CFX removal rate. Figure 3A shows the interaction between the initial CFX concentration and pH. As the initial CFX concentration increased to 10 mg/L, the rate of CFX removal increased and then it slowed down. By reducing the pH from 7.5 to 4.5, the amount of pollutant removal increased. As figure 3B shows, the removal rate increased by increasing starch dose and reducing contact time from 95 min to 45 min.

### Discussion

According to figure 3A, the efficiency of CFX removal increased with decreasing pH. Since surface charge of CFX ions depends on pH value and pKa value of cefixime is 3.73 and 2.1 (The presence of two carboxyl groups), at pH equal to 4.5, CFX is mainly in the form of negative ions<sup>19</sup>. In the present study, the

results showed that with acidification of experimental environment, CFX removal was better performed. Mostafaloo *et al*<sup>6</sup> obtained similar results in the study of CFX removal by BiFeO<sub>3</sub> photo catalyst.

Reducing efficiency of removal at higher CFX concentrations (Figure 3A) can be attributed to saturation and reduction of active sites on the starch surface<sup>6</sup>. In a study by Dehghani *et al*<sup>20</sup> at concentrations of 0.47-0.79  $m\mu$  of sulfadiazine, increasing antibiotic initial concentration reduced efficacy. The same results were obtained in the study of Yoosefian *et al*<sup>21</sup>, and Ouaisa *et al*<sup>22</sup>.

According to the results, with increasing starch dose, the amount of pollutant removal has been increased. Therefore, it can be said increasing adsorption mass is associated with an increase in surface and more active sites for pollutant adsorption<sup>23</sup>. Fakhri<sup>24</sup> in a study on efficacy of CFX removal with MGO nanoparticles, found that increasing the adsorbent dose to 0.45 g/L increases the efficacy of CFX removal and then the upward trend greatly slows down. Therefore, 0.45 g/L was selected as the optimal dose of MGO. Similar conclusions were made in the study of Elmolla<sup>25</sup> and Mostafaloo *et al*<sup>6</sup>.

In this study, the CFX removal rate decreased with increasing time. Rapid adsorption at first contact time is due to large number of active sites on the adsorbent surface, which improves diffusion of cefixime to adsorbent surface<sup>26,27</sup>. This result is in good agreement with Homem *et al*<sup>28</sup> and Mostafaloo *et al*'s<sup>29</sup> investigations.

### Conclusion

In this study, CFX removal was performed using rice starch. The response surface methodology and CCD were used to design, analyze and optimize experiments. ANOVA showed the proposed regression model was consistent with statistical analysis:  $R^2=0.9544$ ,  $R^2$  adjusted=0.9118,  $R^2$  predicted=0.7582.

The optimal values in this study were obtained as follows: maximum CFX removal efficiency, 70.22% with desirability=0.936, optimal pH=4.5, starch dose=225 mg/L, initial CFX concentration=7.9 mg/L and time=95 min.

Due to proper removal of CFX with starch, it is recommended to use the sewage of industries that contain starch in pre-treatment of pharmaceutical industry wastewater.

### Acknowledgement

The authors would like to appreciate the assistance of Qom University of Medical Sciences, Iran.

### Ethical Approval

Authors are aware of, and comply with, best practice in publication ethics specifically with regard to authorship (Avoidance of guest authorship), dual sub-

mission, manipulation of figures, competing interests and compliance with policies on research ethics. Authors adhere to publication requirements that submitted work is original and has not been published elsewhere in any languages.

### Conflict of Interest

The authors declare that they have no conflict of interests.

### References

1. Homem V, Santos L. Degradation and removal methods of antibiotics from aqueous matrices- a review. *J Environ Manage* 2011;92(10):2304-2347.
2. Aydin S, Aydin ME, Beduk F, Ulvi A. Removal of antibiotics from aqueous solution by using magnetic Fe<sub>3</sub>O<sub>4</sub>/red mud-nanoparticles. *Sci Total Environ* 2019;670:539-546.
3. Shaniba C, Akbar M, Ramseena K, Raveendran P, Narayanan BN, Ramakrishnan RM. Sunlight-assisted oxidative degradation of cefixim antibiotic from aqueous medium using TiO<sub>2</sub>/nitrogen doped holey graphene nanocomposite as a high performance photocatalyst. *J Environmental Chemical Engineering* 2018;8(1):102204.
4. Pham TD, Bui TT, Truong TTT, Hoang TH, Le TS, Duong VD, et al. Adsorption characteristics of beta-lactam cefixime onto nanosilica fabricated from rice HUSK with surface modification by polyelectrolyte. *J Mol Liq* 2020;298:111981.
5. Kümmerer K, Alexy R, Hüttig J, Schöll A. standardized tests fail to assess the effects of antibiotics on environmental bacteria. *Water Res* 2004;38(8):2111-2116.
6. Mostafaloo R, Mahmoudian MH, Asadi-Ghalhari M. BiFeO<sub>3</sub>/magnetic nanocomposites for the photocatalytic degradation of cefixime from aqueous solutions under visible light. *J Photochem Photobiol A Chem* 2019;382:111926.
7. Belghadr I, Shams Khorramabadi G, Godini H, Almasian M. The removal of the cefixime antibiotic from aqueous solution using an advanced oxidation process (UV/H<sub>2</sub>O<sub>2</sub>). *Desalination Water Treat* 2015;55(4):1068-1075.
8. Khan MN, Qayum A, Rehman UU, Gulab H, Idrees M. Spectrophotometric method for quantitative determination of cefixim in bulk and pharmaceutical preparation (using ferron complex). *J Appl Spectrosc* 2015;82(4):705-711.
9. Goossens H, Ferech M, Coenen S, Stephens P. Comparison of diseases in the United States and 27 European countries.outpatient systemic antibacterial use in 2004. *Clinical Infectious Diseases* 2007;44(8):1091-1095.
10. Kandhro AA LA, Mahesar SA, Saleem R, Nelofar A, Khan ST, et al. Application of attenuated total reflectance Fourier transform infrared spectroscopy for determination of cefixime in oral pharmaceutical formulations. *Spectrochim Acta A Mol Biomol Spectrosc* 2013;115:51-56.
11. Esmaeili Bidhendi M, Poursorkh Z, Sereshti H, Rashidi Nodeh H, Rezaia S, Afzal Kamboh M. Nano-size bio-

- mass derived from pomegranate peel for enhanced removal of cefixime antibiotic from aqueous media: kinetic, equilibrium and thermodynamic study. *Int J Environ Res Public Health*.2020;17(12):4223.
12. Iram M, Guo C, Guan Y, Ishfaq A, Liu H. Adsorption and magnetic removal of neutral red dye from aqueous solution using Fe<sub>3</sub>O<sub>4</sub> hollow nanospheres. *J Hazard Mater* 2010;181(1-3):1039-1050.
  13. Montanher S, Oliveira E, Rollemberg M. Removal of metal ions from aqueous solutions by sorption onto rice bran. *J Hazard Mater* 2005;117(2-3):207-211.
  14. Wang L, Xie B, Shi J, Xue S, Deng Q, Wei Y, et al. Physicochemical properties and structure of starches from Chinese rice cultivars. *Food Hydrocolloids* 2010;24(2-3):208-216.
  15. Dehghani S, Jonidi Jafari A, Farzadkia M, Gholami M. Sulfonamide antibiotic reduction in aquatic environment by application of fenton oxidation process. *Iranian J Environ Health Sci Eng* 2013;10(1):29.
  16. Kobya M, Demirbas E, Gebologlu U, Oncel M, Yildirim Y. Optimization of arsenic removal from drinking water by electrocoagulation batch process using response surface methodology. *Desalination Water Treat* 2013;51(34-36):6676-6687.
  17. Mostafaloo R, Asadi-Ghalhari M, Izanloo H, Zayadi A. Photocatalytic degradation of ciprofloxacin antibiotic from aqueous solution by BiFeO<sub>3</sub> nanocomposites using response surface methodology. *Global J Environmental Science Management* 2020;6(2):191-202.
  18. Karimi R, Yousefi F, Ghaedi M, Dashtian K. Back propagation artificial neural network and central composite design modeling of operational parameter impact for sunset yellow and azur (II) adsorption onto MWCNT and MWCNT-Pd-NPs: Isotherm and kinetic study. *Chemometr Intell Lab Syst* 2016;159:127-137.
  19. Anaconda J, Estacio J. Synthesis and antibacterial activity of cefixime metal complexes. *Transition Metal Chemistry* 2006;31(2):227-231.
  20. Dehghani S JJA, Farzadkia M, Gholami M. [Investigation of the efficiency of Fenton's advanced oxidation process in sulfadiazine antibiotic removal from aqueous solutions]. *Arak Med Univ J* 2012;15(66):19-290. Persian.
  21. Yoosefian M, Ahmadzadeh S, Aghasi M, Dolatabadi M. Optimization of electrocoagulation process for efficient removal of ciprofloxacin antibiotic using iron electrode: kinetic and isotherm studies of adsorption. *J Mol Liq* 2016;225:544-553.
  22. Ouaisa YA, Chabani M, Amrane A, Bensmaili A. Removal of tetracycline by electrocoagulation: Kinetic and isotherm modeling through adsorption. *J Environ Chem Eng* 2014;2(1):177-184.
  23. Maghsoudi M, Ghaedi M, Zinali A, Ghaedi A, Habibi M. Artificial neural network (ANN) method for modeling of sunset yellow dye adsorption using zinc oxide nanorods loaded on activated carbon: Kinetic and isotherm study. *Spectrochim Acta A Mol Biomol Spectrosc* 2015;134:1-9.
  24. Fakhri A, Adami S. Adsorption and thermodynamic study of Cephalosporins antibiotics from aqueous solution onto MgO nanoparticles. *J Taiwan Inst Chem Eng* 2014;45(3):1001-1006.
  25. Elmolla ES, Chaudhuri M. Photocatalytic degradation of amoxicillin, ampicillin and cloxacillin antibiotics in aqueous solution using UV/TiO<sub>2</sub> and UV/H<sub>2</sub>O<sub>2</sub>/TiO<sub>2</sub> photocatalysis. *Desalination* 2010;252(1-3):46-52.
  26. Roy P, Dey U, Chattoraj S, Mukhopadhyay D, Mondal NK. Modeling of the adsorptive removal of arsenic (III) using plant biomass: a bioremediation approach. *Appl Water Sci* 2017;7(3):1307-1321.
  27. Ghaedi M. Comparison of cadmium hydroxide nanowires and silver nanoparticles loaded on activated carbon as new adsorbents for efficient removal of Sunset yellow: Kinetics and equilibrium study. *Spectrochim Acta A Mol Biomol Spectrosc* 2012;94:346-351.
  28. Homem V, Alves A, Santos L. Amoxicillin degradation at ppb levels by Fenton's oxidation using design of experiments. *Sci Total Environ* 2010;408(24):6272-6280.
  29. Mostafaloo R, Yari AR, Mohammadi MJ, Khaniabadi YO, Asadi-Ghalhari M. Optimization of the electrocoagulation process on the effectiveness of removal of Cefixime antibiotic from aqueous solutions. *Desalination Water Treat* 2019;144:138-144.