

***Typha domingensis* leaf powder for decontamination of aluminium, iron, zinc and lead: Biosorption kinetics and equilibrium modeling**

¹N. T. Abdel-Ghani; ²A. K. Hegazy; ³*G. A. El-Chaghaby

¹Chemistry Department, Faculty of Science, Cairo University, Giza, Egypt

²Botany Department, Faculty of Science, Cairo University, Giza, Egypt

³Agriculture Research Center, Giza, Egypt

Received 22 December 2008; revised 14 January 2009; accepted 19 February 2009; available online 1 March 2009

ABSTRACT: The present study explores the effectiveness of *Typha domingensis* leaf powder for simultaneous removal of aluminium, iron, zinc and lead ions from aqueous solution. Batch experiments were carried out in laboratory at room temperature and at initial ions concentrations simulating the concentrations of these cations in real wastewater samples. The sorption process was examined applying the first and second order kinetic mechanisms. The results were best described by the second order rate kinetics. The applicability of the three equilibrium isotherm models was investigated. The obtained data follow the three investigated isothermal models in the following order: Langmuir > Freundlich > Temkin, for all the studied metal ions. The infrared spectra of native and exhausted *Typha* leaf powder confirmed ions-biomass interactions responsible for sorption. The results showed that *Typha domingensis* leaf powder can easily be envisaged as a new low cost natural biosorbent for metal clean up operations in aquatic systems.

Keywords: Batch experiments, contact time, infrared spectra, isotherm models

INTRODUCTION

Wastewater discharged by industrial activities has been a matter of worldwide concern over the last few decades posing serious environmental concerns. Metal ions represent a major class of pollutants introduced into aquatic system significantly as a result of various industrial and agricultural operations (Chandra *et al.*, 2003). Different countries are devoting concentrated efforts for the treatment and removal of heavy metals in order to combat this problem. The commonly used procedures for removing metal ions from aquatic ecosystems include chemical precipitation, reverse osmosis and solvent extraction (Bayat, 2002; Gupta and Torres, 1998). However, these techniques have certain disadvantages such as incomplete metal removal, high reagent and energy requirements, generation of toxic sludge or other waste products that require disposal (Chandra *et al.*, 2003). The search for alternate and innovative treatment techniques has focused attention on the use of biological materials for heavy metal removal and recovery technologies (biosorption) and has gained important credibility during recent years because of the good performance and low cost of this complexing materials (Chandra *et al.*, 2003). Different types of

biosorbents were investigated for metals removal from wastewater (Abdel-Ghani and El-Chaghaby, 2008; Abdel-Ghani *et al.*, 2008; Horsfall and Abia, 2003; Singh *et al.*, 2005; Zafar *et al.*, 2007). Among many biosorbents naturally available plant biomass can be used as an inexpensive method for metal ions removal from synthetic solutions and industrial effluents (Manal, 2007). In the present investigation, the potential of *Typha domingensis* leaf powder has been assessed for the removal of metal ions such as aluminium, iron, zinc and lead. *Typha domingensis* was selected as it is a widespread and dominant plant in many aquatic systems. The species also grows in industrial wastewater lakes which indicate that the plant is resistant to pollution. The effects of various parameters affecting the removal of metal ions by biosorption onto *Typha domingensis* leaf powder were studied and the results are discussed.

The present study was performed during 2008 at the agriculture research center in Egypt.

MATERIALS AND METHODS

Biosorbent preparation and characterization

Fresh biomass of *Typha domingensis* were collected from Sadat City, Menoufia governorate, Egypt. Before use, the biomass was washed with deionized water to remove dirt. The biomass was then air dried for 10 days

*Corresponding Author Email: ghadiraly@yahoo.com
Tel./Fax: +0101 235 184

followed by drying in an oven at 70 °C for 24 h. Subsequently, it was ground. Elemental analysis of the prepared biosorbent was estimated. The amounts of hemicellulose, lignin and cellulose in the biosorbent material were determined. Fourier transformation infrared (FTIR) of the biosorbent was also obtained.

Determination of Al, Fe, Zn and Pb

Al (III), Fe(III), Zn(II) and Pb(II) were simultaneously determined by using a Perkin Elmer inductively coupled plasma optical emission spectrometer (ICP-OES), model Optima 2000DV.

Batch experiments

Batch sorption experiments were performed at a constant temperature of 25 °C in a Shaking water bath at 100 rpm using 125 mL capped conical flasks. Al (III), Fe (III), Zn (II) and Pb (II) working solutions were prepared in deionized water using aluminium, iron, zinc and lead 1000 ppm stock solutions (Merck). In all sets of experiments, 0.5 g of *Typha* leaves powder was thoroughly mixed into 50 ml mixed Al, Fe, Zn and Pb ions solutions. In order to investigate the effect of contact time on metals removal by adsorption onto *Typha* biomass a known weight (0.5 g) of biomass was added to a mixed metal ion solution containing 7 ppm (Al), 10 ppm (Fe), 4 ppm (Zn) and 1 ppm (Pb) simulating the concentrations of these cations in real wastewater samples collected from industrial wastewater lake at Sadat City, Egypt. The pH was measured and was found to be 2.5. The solutions were shaken well in a shaking water bath at 100 rpm at room temperature for 30, 60, 90, 120, 150 and 240 min. The slurry, after equilibration, was filtered through Whatman 41 filter paper. The clear filtrate was analyzed for metal concentration using a Perkin Elmer inductively coupled plasma- optical emission spectrometer ICP-OES Optima 2000 DV. The amount of metal taken up by plant biomass was calculated as the difference between the initial and final concentrations of metal after adsorption in the aqueous solution. The removal efficiency was calculated according to the following equation:

$$\text{Removal efficiency (R)} = (C_0 - C_f) \times 100 / C_0 \quad (1)$$

Where, C_0 and C_f are the initial and final ions concentrations, respectively. In order to investigate the mechanism of biosorption, the pseudo first order and pseudo second order equations were used to test experimental data. The first order rate expression of

(Lagergren, 1898) based on solid capacity is generally expressed as follows:

$$dQ/dt = k_{1,ads} (Q_{eq} - Q) \quad (2)$$

The pseudo second-order equation is also based on the sorption capacity of the solid phase (Ho and McKay, 1999), contrary to the other mechanism being the rate controlling step. If the rate of sorption is a second order mechanism, the pseudo second order chemisorption kinetic rate equation is expressed as:

$$dQ/dt = k_{2,ads} (Q_{eq} - Q)^2 \quad (3)$$

The experimental amount of metal ion biosorption (Q_{eq}) per gram of biomass was calculated by the following equation:

$$Q_{eq} = V (C_0 - C_f) / m_s \quad (4)$$

Where, C_0 (mg /L) is the initial concentration, C_f (mg /L) is the final or equilibrium concentration, V (L) is the volume of solution and m_s (g) is the biosorbent mass (dry base). Equilibrium data, which are commonly known as adsorption isotherms, are basic requirements for the design of adsorption systems. This set of experiments was performed at different initial ions concentrations. Classical adsorption models (Langmuir, Freundlich and Temkin) were used to describe the equilibrium between adsorbed metal ions on the *Typha* biomass (Q_{eq}) and metal ions in solution (C_{eq}) at a constant temperature.

RESULTS AND DISCUSSION

Biosorbent characterization

Biosorbent composition

Typha domingensis leaf powder subjected to elemental analysis showed composition of carbon, hydrogen, nitrogen and sulphur as 43.35, 10.06, 2.49 and 1.21 %, respectively. The leaf powder was also found to contain 31.17, 34.09 and 4.16 % hemicellulose, cellulose and lignin, respectively.

Infrared spectra of the biosorbent

The FTIR spectra of fresh-dried (A) metal-loaded (B) *Typha* biomass are shown in Fig. 1. The spectra display a number of absorption peaks, indicating the complex nature of the examined biomass. Inspection of the spectra reveals the presence of a broad peak at wave number 3400.85/cm representing surface bonded O-H groups or stretching N-H groups which was shifted to 3419.17/cm after adsorption; The strong absorption

peak at 2921.63/cm can be assigned to –CH stretching, and those at 1735.62/cm and 1641.13/cm assigned to C=O stretching and CO chelate stretching of carboxyl groups, respectively. These peaks also were shifted to 2919.7/cm, 1736.58/cm and 1648.84/cm. The absorption peaks at 1424.1/cm and 1246.75/cm can be attributed to the presence of (CH₂ and CH₃) and –C–O stretching, respectively. After adsorption these peaks were shifted to 1380.78 and 1247.72/cm. Another peak observed at the biomass surface before adsorption appeared at 1059.69/cm and was shifted to 1057.76/cm after adsorption. The peak observed at 603.61/cm that was shifted to 612.288/cm after adsorption could be assigned to out of plane O-H. The shifts in the absorption peaks generally observed indicate the existence of a metal binding process taking place on the surface of *Typha* biomass (Abdel-Ghani and El-Chaghaby, 2007; Abdel-Ghani *et al.*, 2007; Horsfall *et al.*, 2006; Ilharco *et al.*, 1997).

Effect of Contact time and Biosorption kinetics

The contact time and kinetic parameters are necessary to determine the best operational conditions in a continuous process of ions removal. Many previous authors have investigated the effect of contact time on biosorption process, as well as the biosorption kinetics (El-Ashtoukhy *et al.*, 2008; Pandey *et al.*, 2008; Selatnia *et al.*, 2004; Al-Anber and Matouq, 2008). As shown in Fig. 2, the removal of Al (III), Fe (III), Zn (II) and Pb(II) by adsorption onto *Typha* biomass increases with contact time and attains saturation in about 120 min. Basically, the removal rate of sorbate is rapid but it

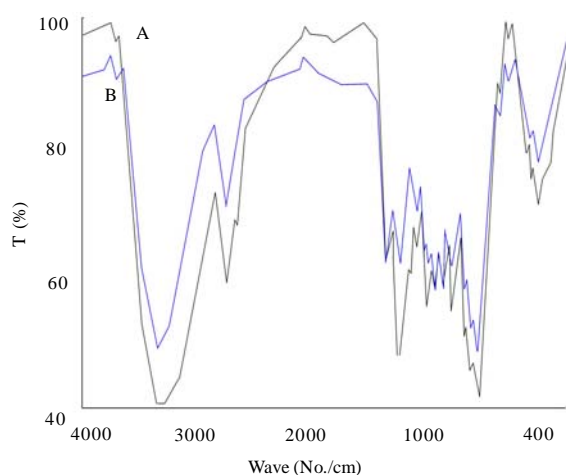


Fig. 1: FTIR spectra of fresh-dried (A) metal-loaded (B) *Typha* biomass

gradually decreases with time until it reaches equilibrium. The rate of percent metal removal is higher in the beginning due to a larger surface area of the adsorbent being available for the adsorption of the metals (El-Ashtoukhy *et al.*, 2008). It is also relevant to point out that, since active sorption sites in a system have a fixed number and each active site can adsorb only one ion in a monolayer (Langmuir, 1918), the metal uptake by the sorbent surface will be rapid initially, slowing down as the competition for decreasing availability of active sites intensifies by the metal ions remaining in solution. The rate of metal removal is of greatest significance for developing a natural adsorbent-based water-treatment technology. In order to investigate the mechanism of biosorption, kinetic models were used to test experimental data. The kinetic models included the pseudo first order and pseudo second order equations.

After integration and applying boundary conditions, $t = 0$ to $t = t$ and $Q = 0$ to $Q = Q_{eq}$; the integrated form of Eq. 2 becomes:

$$\ln(Q_{eq} - Q) = \ln Q_{eq} - k_{1,ads} t \tag{5}$$

A straight line of $\ln(Q_{eq} - Q)$ versus t suggests the applicability of the pseudo first-order kinetic model (Lagergren, 1898; McKay *et al.*, 1999; Ho and McKay, 1999). The rate constant $k_{1,ads}$ and the equilibrium amount of metal ion, Q_{eq} , of Lagergren pseudo first order model can be, respectively, obtained from the slope and intercept of plot. In order to fit Eq. 5 to experimental data, the equilibrium sorption capacity, Q_{eq} , must be known. In many cases, Q_{eq} is unknown and as adsorption tends to become unmeasurably slow, the amount sorbed is still significantly smaller than the equilibrium amount. For this reason, it is necessary to obtain the real equilibrium sorption capacity, Q_{eq} , by extrapolating the experimental data to $t = 1$ or using a trial and error method. Furthermore, in most cases, the first order equation of Lagergren does not fit well for the whole range of contact time and is generally applicable over the initial 20–30 min of the sorption process.

For the boundary conditions, $t = 0$ to $t = t$ and $Q = 0$ to $Q = Q_{eq}$, the integrated form of Eq. 3 becomes Eq. 6:

$$1/Q_{eq} - Q = 1/Q_{eq} + k_{2,ads} t \tag{6}$$

Eq. 6 can be rearranged to a linear form:

Typha domingensis leaf powder for wastewater decontamination

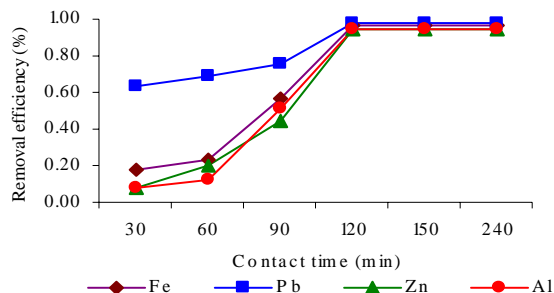


Fig. 2: Effect of contact time on the removal of Al (III), Fe (III), Zn (II) and Pb(II) by adsorption onto *Typha* biomass; pH:2.5; biomass weight: 0.5 g/50 mL; shaking speed: 100 rpm; temperature: 250 °C

$$(t / Q) = 1/k_{2,ads} Q_{eq}^2 + (1/ Q_{eq}) t \quad (7)$$

If second-order kinetics are applicable, the plot of t/Q against t of Eq. 6 should give a linear relationship, from which Q_{eq} and $k_{2,ads}$ can be determined from the slope and intercept of the plot, respectively and there is no need to know any parameter before hand (Ho and McKay, 1999). The pseudo first-order and pseudo second-order kinetic models were applied to the experimental data. The first-order and the second-order sorption rate constants for Al^{3+} , Fe^{3+} , Zn^{2+} and Pb^{2+} as well as the correlation coefficients obtained for each model are presented in Table 1. It can be concluded that the obtained correlation coefficients for the second-order kinetic model were greater than those obtained for the first-order kinetic model. The theoretical Q_{eq} values were close to the experimental Q_{eq} values in the case of pseudo second-order kinetics. These suggest that the sorption system is the second-order model, based on the assumption that the rate determining step may be the biosorption, provides the best correlation of the data.

Equilibrium modeling

The Langmuir equation is valid for monolayer sorption onto a surface with a finite number of identical sites and is given by Eq 8:

$$C_{eq} / Q_{eq} = 1/b Q_{max} + C_{eq} / Q_{max} \quad (8)$$

Where, Q_{max} (mg/g) is the maximum amount of metal ion per unit weight of *Typha* leaf powder to form a complete monolayer on the surface bound at high C_{eq} , and b (L/mg) is a constant related to the affinity of the biomass surface binding sites when the surface is fully covered with metal ions and assists in the comparison of adsorption performance, particularly in cases where the sorbent did

not reach its full saturation. The Q_{max} represents a practical limiting adsorption capacity in experiments. The Q_{max} and b can be determined from the linear plot of C_{eq} / Q_{eq} versus C_{eq} (Donmez *et al.*, 1999; Matheickal *et al.*, 1999; McKay *et al.*, 1999; Veglio and Beolchini, 1997).

Freundlich expression is an empirical equation based on sorption on a heterogeneous surface (Freundlich, 1906). The empirical Freundlich equation based on sorption on a heterogeneous surface is given below by Eq. 9:

$$\log Q = \log K_f + 1/n \log C_{eq} \quad (9)$$

Where, (K_f) and (n) are the Freundlich constants characteristic of the system and are indicators of adsorption capacity and adsorption intensity, respectively. The derivation of the Temkin isotherm (Kolasniski, 2001) assumes that the fall in the heat of sorption is linear rather than logarithmic, as implied in the Freundlich equation.

The Temkin isotherm is generally applied in the following form:

$$Q_{eq} = a_t + b_t \ln C_{eq} \quad (10)$$

Where, b_t and a_t represent isotherm constants related to the adsorption capacity of the adsorbent. The linearized Langmuir, Freundlich and Temkin adsorption isotherms of Al^{3+} , Fe^{3+} , Zn^{2+} and Pb^{2+} ions were obtained at the temperature of 25 °C. The three models adsorption constants evaluated from the isotherms with their corresponding correlation coefficients are presented in Table 2. As seen from Table 2, very high regression correlation coefficients ($R^2 > 0.9$) were found for the Langmuir model, whereas high regression correlation coefficient for Freundlich model was found only in case of Pb^{2+} . According to these results, the Langmuir model is more suitable for describing the biosorption equilibrium of Al^{3+} , Fe^{3+} , Zn^{2+} and Pb^{2+} by *Typha* biomass in the studied concentration range. An adsorption isotherm is characterized by certain constants; the values of which express the surface properties and affinity of the sorbent and can also be used to find the biosorptive capacity of biomass. The values of R^2 are regarded as a measure of the goodness-of-fit of experimental data on the isotherm models (El-Ashtoukhy *et al.*, 2008). The applicability of the three isotherm models for the present data follows approximately the order: Langmuir > Freundlich > Temkin, for all the studied metal ions. The essential characteristic of the Langmuir isotherm may be expressed in terms of dimensionless separation parameter R_L , which is

Table 1: The first-order and the second-order sorption rate constants for Al³⁺, Fe³⁺, Zn²⁺ and Pb²⁺

	Pseudo first order			Pseudo second order			Q _{eq} experimental (mg/g)
	K _{1,ad} /min	Q _{eq} (mg/g)	R ²	K _{2,ad} g/mg min	Q _{eq} (mg/g)	R ²	
Al ³⁺	0.0048	0.98	0.918	0.0622	0.677	1	0.67875
Fe ³⁺	0.0098	1.04	0.8932	0.0465	0.99	0.8948	0.9291
Zn ²⁺	0.0199	0.76	0.8667	0.0413	0.29	1	0.2945
Pb ²⁺	0.037	0.40	0.7005	0.267	0.132	0.9825	0.1167

Temp: 25 °C; agitation rate: 100 rpm; biomass concentration: 0.5 g/50 mL; pH: 2.5; initial ions concentrations: Al³⁺ (7 mg/L), Fe³⁺ (10 mg/L), Zn²⁺ (4 mg/L) and Pb²⁺(1 mg/L)

Table 2: Sorption isotherm coefficients of Langmuir, Freundlich and Temkin models applied to Al³⁺, Fe³⁺, Zn²⁺ and Pb²⁺ biosorption by *Typha* biomass

	Langmuir			Freundlich			Temkin		
	Q _{max} , mg/g)	b (L/mg)	R ²	K _f	n	R ²	a _t	b _t	R ²
Al ³⁺	0.34825	0.287691	0.9654	0.1538	3.82995	0.5053	1.0012	0.815	0.4584
Fe ³⁺	2.111932	2.107927	0.9778	1.073	5.77034	0.8883	2.3964	11.091	0.8296
Zn ²⁺	0.201349	5.412609	0.9443	0.05516	2.282584	0.8091	0.654	0.0523	0.8028
Pb ²⁺	0.654579	1.200842	0.9951	0.32047	3.522367	0.9437	4.0349	0.8814	0.8608

Temp: 25 °C; agitation rate: 100 rpm; biomass concentration: 0.5g/50mL; pH: 2.5

Table 3: Different values of R_L for the different concentrations of Al³⁺, Fe³⁺, Zn²⁺ and Pb²⁺

Al ³⁺ concentration (mg/L)	R _L	Fe ³⁺ concentration (mg/L)	R _L	Zn ²⁺ concentration (mg/L)	R _L	Pb ²⁺ concentration (mg/L)	R _L
7	0.331	10	0.045	4	0.044	1	0.454
14	0.198	20	0.023	8	0.022	5	0.142
28	0.110	40	0.011	16	0.011	15	0.052
49	0.066	50	0.009	24	0.007	25	0.032
56	0.058	60	0.007	32	0.005	30	0.027

indicative of the isotherm shape that predicts whether an adsorption system is favourable or unfavourable.

$$R_L = 1 / (1 + b C_0)$$

R_L is defined as (Webi and Chakravort, 1974):

Where, b is the Langmuir constant. Values of R_L for Al³⁺, Fe³⁺, Zn²⁺ and Pb²⁺ are given in Table 3. The adsorption process as a function of R_L may be described as (Chakravarty *et al.*, 2009): Unfavorable when R_L > 1; Linear when R_L = 1; Favorable when 0 < R_L < 1; Irreversible when R_L = 0. In all cases, the R_L values for the present experimental data fall between zero and one, which is an indication of the favourable adsorption of the studied metal ions on the biosorbent (El-Ashtoukhy *et al.*, 2008).

CONCLUSION

Typha domingensis leaf powder can be used to simultaneously remove aluminium, iron, zinc and lead ions from aqueous solutions. It was noted that the removal process increases with contact time and attains saturation in about 120 min. Metal/biomass isotherms were developed and analyzed according to three isotherm equations. The experimental results were analyzed using the Langmuir, Freundlich and Temkin

equations and the data were found to fit the examined models in the order: Langmuir > Freundlich > Temkin, for all the studied metal ions. Adsorption kinetics follows pseudo-second order kinetic model. FTIR investigation of the biomass before and after adsorption showed many changes and shifts in some peaks indicating that there is a metal binding process taking place on the surface of *Typha* biomass. *Typha domingensis* biomass was successfully used for Al³⁺, Fe³⁺, Zn²⁺ and Pb²⁺ removal from simulated wastewater. More parameters could also be examined to assess the possibility of the use of biomass for technological applications.

REFERENCES

- Abdel-Ghani N. T.; El-Chaghaby, G. A., (2007). Influence of operating conditions on the removal of Cu, Zn, Cd and Pb ions from wastewater by adsorption. Int. J. Environ. Sci. Tech., 4 (4), 451-456 (6 pages).
- Abdel-Ghani, N. T.; El-Chaghaby, G. A., (2008). The use low cost, environmental friendly materials for the removal of heavy metals from aqueous solutions. Curr. World Environ., 3 (1), 31-38 (8 pages).
- Abdel-Ghani, N. T.; Hefny, M.; El-Chaghaby, G. A. F., (2007). Removal of lead from aqueous solution using low-cost abundantly available adsorbents. Int. J. Environ. Sci. Tech., 4 (1), 67-73 (7 pages).

- Abdel-Ghani, N. T.; Hefny, M.; El-Chaghaby, G. A., (2008). Removal of metal ions from synthetic wastewater by adsorption onto Eucalyptus camaldulensis tree leaves. J. Chilean Chem. Soc., 53 (3), 1585-1587 (3 pages).
- Al-Anber, Z. A.; Matouq, M. A. D., (2008). Batch adsorption of cadmium ions from aqueous solution by means of olive cake. J. Hazard. Mater., 151 (1), 194-201 (8 pages).
- Bayat, B., (2002). Combined removal of zinc (II) and cadmium (II) from aqueous solutions by adsorption onto high-calcium Turkish fly ash., Water Air Soil Poll. 136 (1-4), 69-92 (24 pages).
- Chakravarty, S.; Pimple, S.; Hema, S.; Chaturvedi, T.; Singh, S.; Gupta, K. K., (2009). Removal of copper from aqueous solution using newspaper pulp as an adsorbent., J. Hazard. Mater., 159 (2-3), 396-403 (8 pages).
- Chandra S., K.; Kamala, C. T.; Chary, N. S.; Anjaneyulu, Y., (2003). Removal of heavy metals using a plant biomass with reference to environmental control. Int. J. Miner. Proc., 68 (1-2), 37-45 (9 pages).
- Donmez, G. C.; Aksu, Z.; Ozturk, A.; Kutsal, T., (1999). A comparative study on heavy metal biosorption characteristics of some algae. Proc. Bioch., 34 (9), 885-892 (8 pages).
- Ekmekyapar, F.; Aslan, A.; Kemal Bayhan, Y.; Cakici A., (2006). Biosorption of copper (II) by nonliving lichen biomass of Cladonia rangiformis hoffm. J. Hazard. Mater., B137, 293-298 (6 pages).
- El-Ashtoukhy, E. S. Z.; Amin, N. K.; Abdelwahab, O., (2008). Removal of lead (II) and copper (II) from aqueous solution using pomegranate peel as a new adsorbent. Desalination, 223 (1-3), 162-173 (12 pages).
- Freundlich, H. M. F., (1906). Über die adsorption in lösungen. Zeitschrift für Physikalische Chemie. 57, 385-470 (85 pages).
- Huama'n Pino, G., Souza de Mesquita, L. M., Torem M. L.; Saavedra Pinto, G. A., (2006). Biosorption of cadmium by green coconut shell powder. Mine. Eng., 19 (5), 380-387 (8 pages).
- Gupta, G.; Torres, N., (1998). Use of fly ash in reducing toxicity of and heavy metals in wastewater effluent. J. Hazard. Mater. 57 (1), 243-248 (6 pages).
- Ho, Y. S.; McKay, G., (1999). Pseudo-second order model for sorption processes. Proc. Biochem., 34(5):451-465 (5 pages).
- Horsfall, M. Jr.; Abia, A. A., (2003). Sorption of Cd(II) and Zn(II) ions from aqueous solutions by cassava waste biomass (Manihot sculenta Cranz). Water Res., 37 (20), 4913-4923 (11 pages).
- Horsfall, M. J.; Ogban, F. E.; Akporhonor, E. E., (2006). Recovery of lead and cadmium ions from metal-loaded biomass of wild cocoyam (*Caladium bicolor*) using acidic, basic and neutral eluent solutions. Electron. J. Biotech., 9 (2), 152-156 (5 pages).
- Ilharco, L. M.; Garcia, A. R.; Lopes da Silva, J.; Vieira Ferreira, L. F., (1997). Infrared approach to the study of adsorption on cellulose: Influence of cellulose crystallinity on the adsorption of benzophenone. Langmuir, 13 (15), 4126-4132 (7 pages).
- Kolasniski, K. W., (2001). Surface Science. Wiley, Chister, UK.
- Lagergren S., (1898). Zur theorie der sogenannten adsorption gelöster stoffe. Kungliga Svenska Vetenskapsakademiens. Handlingar, 24 (4), 1-39 (39 pages).
- Langmuir, I., (1918). The adsorption of gases on plane surfaces of glass, mica and platinum. J. Am. Chem. Soc., 40 (9), 1361-1368 (8 pages).
- Manal, F., (2007). Biosorption of cadmium and lead by phragmites Australis L. biomass using factorial experiment design. Global j. biotech. biochem., 2 (1), 10-20 (11 pages).
- Matheickal, J. T.; Yu, Q.; Yin, P.; Kaewsarn, P., (1999). Heavy metal uptake capacities of common marine macro algal biomass. Water Res., 33 (6), 1534-1537 (4 pages).
- McKay, G.; Ho, Y. S.; Ng, J. C. Y., (1999). Biosorption of copper from wastewaters: A review. Separ. Purif. Method., 28 (1), 87-125 (39 pages).
- Pandey, P. K.; Verma, Y.; Choubey, Sh., Pandey, M.; Chandrasekhar, K., (2008). Biosorptive removal of cadmium from contaminated groundwater and industrial effluents. Bioresource Tech., 99 (10), 4420-4427 (8 pages).
- Selatnia, A.; Boukazoula, A.; Kechid, N.; Bakhti, M. Z.; Chergui, A., (2004). Biosorption of Fe³⁺ from aqueous solution by a bacterial dead *Streptomyces rimosus* biomass. Process Biochem., 39 (11), 1643-1651 (9 pages).
- Singh, K. K.; Rastogi, R.; Hasan, S. H., (2005). Removal of cadmium from wastewater using agricultural waste rice polish. J. Hazard. Mater., 121 (1-3), 51-58 (8 pages).
- Veglio, F.; Beolchini, F., (1997). Removal of metals by biosorption: a review. Hydrometallurgy, 44 (3), 301-316 (16 pages).
- Webi T. W.; Chakravort, R. K., (1974). Pore and solid diffusion models for fixed bed adsorbents. J. Am. Inst. Chem. Eng., 20 (2), 228-238 (11 pages).
- Zafar, M. N.; Nadeem, R.; Hanif, M. A., (2007). Biosorption of nickel from protonated rice bran. J. Hazard. Mater., 143 (1-2), 478-485 (8 pages).

AUTHOR (S) BIOSKETCHES

Abdel Ghani, N. T., Professor of analytical and inorganic chemistry in Chemistry Department, Faculty of Science, Cairo University, Egypt. Email: noureta2002@yahoo.com

Hegazy, A. K., is a professor of Conservation and Applied Ecology in Botany Department, Faculty of Science, Cairo University, Egypt. Email: akhegazy@yahoo.com

El-Chaghaby, G. A., M.Sc., is an Assistant researcher in Agriculture Research Center, Egypt; Ph.D. student in Chemistry Department, faculty of science, Cairo University, Egypt. Email: ghadiraly@yahoo.com

This article should be referenced as follows:

Abdel Ghani, N. T.; Hegazy, A. K.; El-Chaghaby, G. A., (2009). *Typha domingensis* leaf powder for decontamination of aluminium, iron, zinc and lead: Biosorption kinetics and equilibrium modeling. Int. J. Environ. Sci. Tech., 6 (2), 243-248.