

# Performance of multi-wall carbon nanotubes modified by chitosan and EDTA in the removal of Direct Blue 15 dye from aqueous solutions

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## ABSTRACT

Dyes are considered to be important pollutants in different industries (e.g., textile) and contain various organic materials with complicated structures. These compounds are mainly carcinogenic and toxic and less biodegradable when discharged into the environment. The present study aimed to investigate the performance of multi-wall carbon nanotubes (MWCNTs) modified by chitosan and EDTA in the removal of Direct Blue 15 dye from aqueous solutions. MWCNT was synthesized and characterized using the BET, zeta potential, DLS, SEM, and FTIR techniques. The effects of key parameters including pH, adsorbent dose, initial dye concentration, and contact time were also evaluated. The experimental data of the adsorption process were analyzed using the Langmuir and Freundlich models. With the increased contact time, the removal efficiency of the dye improved, while the increased pH, initial dye concentration, and adsorbent dose led to the reduced dye removal efficiency. With the optimum values of pH (=7), contact time (60 min), adsorbent dose (0.5 g/L), and initial dye concentration (60 mg/L), the maximum adsorption capacity was determined to be 114.42 mg/g. According to the results, the adsorption process using the modified MWCNT followed the Langmuir model and pseudo-second-order kinetics.

**Keywords:** Parameter optimization, Adsorption, Modification, Chitosan, EDTA

## Introduction

Environmental pollution due to various sources has been a pressing issue in recent years. The textile industry is a basic industry and a major source of environmental pollution. Approximately 700,000 tons of dyes are synthesized worldwide each year, 50% of

which include azo dyes.<sup>1</sup> Azo dyes contain one or more azo bonds (-N=N-) in their structure and are considered to be the most significant synthetic dyes that are extensively produced and have numerous applications in various industries, especially the textile and dyeing industries.<sup>2</sup> Approximately 15% of the annual production of synthetic dyes enters textile effluents during the dyeing process. The discharge of dye-laden wastewater into the environment hinders light from penetrating the aquatic ecosystems, thereby slowing down the photosynthesis process.<sup>3</sup> Moreover, some azo dyes and their impurities (e.g., amines) are

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toxic and even carcinogenic to various organisms.<sup>4</sup>

With the improvement of the textile dyeing industry, new dyes are injected into the market every day with better dyeing qualities. Due to the structural stability of these dyes, such improvement has caused these materials to resist biological and even chemical decomposition. Consequently, several researchers have attempted to find new methods to eliminate dyes from textile wastewater.<sup>5</sup>

The treatment of the wastewater containing dyes is regularly performed using physicochemical and biological techniques.<sup>6</sup> The conventional methods of textile wastewater treatment involve several steps, which are subjected to the product manufacturing process specifications. The treatments often comprise of coagulation and flocculation processes, which could be used individually or in combination with biological filtration processes to remove suspended solids and organic matters and advance dye removal from textile wastewater. One of the drawbacks of such treatments is the formation of high quantities of sludge that is brimmed with dyes and pollutants, which is considered highly challenging to meet environmental standards.<sup>4</sup> Furthermore, biological treatment,<sup>7</sup> ozonation,<sup>8</sup> chemical oxidation and photocatalytic processes,<sup>9</sup> membrane processes,<sup>10</sup> nanoparticles,<sup>11</sup> and photochemical and enzymatic treatment processes<sup>12</sup> have been employed to degrade dyes from textile effluents, and each process has specific advantages and disadvantages.<sup>13</sup>

Adsorption processes are among the practical methods of dye removal from textile wastewater.<sup>14</sup> These processes are highly efficient in the removal of dyes, odors, and organic matters from wastewater.<sup>15</sup> Additionally, the adsorption process is a reliable purification technique for the reduction of investment costs owing to its ease of design and operation and no sensitivity to toxic compounds.<sup>16</sup> Some of the adsorbents used in

the treatment of dyed wastewater include activated carbon,<sup>17</sup> porous carbon,<sup>18</sup> clay,<sup>19</sup> biopolymers (e.g., chitosan),<sup>20</sup> and agricultural waste.<sup>21</sup> In this method, the dye compounds in wastewater are readily transformed to the solid phase, and the adsorbent could be restored and reused.<sup>22</sup> Recently, carbon allotropies such as carbon nanotubes (CNTs) and graphene oxide have been broadly used for the removal of various pollutants (especially dyes) due to their high surface-to-volume proportion and high adsorption capacity.<sup>23</sup> CNTs are composed of one-atom-thick, hollow, cylindrical carbon plates, and their unique properties (e.g., high cross-sectional area, porous and layered structure) make CNTs an effective and practical adsorbent for the removal of various pollutants, especially organic pollutants.<sup>24</sup>

As a non-toxic, biodegradable, and environmentally friendly natural biopolymer, chitosan is considered to be a proper alternative to bleaching as a sorbent in the textile industry. Chitosan is extensively utilized as a biosorbent owing to its chelating behavior and cost-efficiency compared to commercial adsorbents.<sup>25</sup> In addition, chitosan could enhance the efficiency of the adsorption process by modifying or functionalizing the surface of carbon nanotubes through the improvement of interfacial interactions, including chemical, physical, and mechanical bonds.<sup>26</sup>

The present study aimed to investigate the efficiency of the multi-wall carbon nanotubes (MWCNTs) modified with EDTA and chitosan in the removal of Direct Blue 15 dye from aquatic solutions.

## Materials and Methods

### Materials and reagents

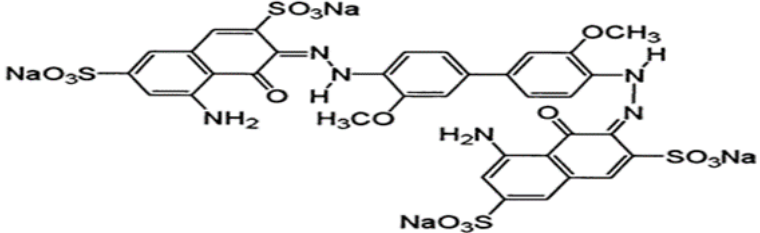
This experimental-laboratory study was conducted to define the efficiency of the EDTA-chitosan-modified MWCNT composites in a batch mode in the removal of Direct Blue 15 dye from aquatic environments. The MWCNTs were purchased from the



Research Institute of Petroleum Industry, Iran, and Direct Blue 15 was provided by Alvan Sabet Co. in Hamadan, Iran (Table 1). The maximum adsorption wavelength of the Direct Blue 15 dye was determined using a UV/Vis spectrophotometer (model: DR-5000-HACH). To determine the location of the functional groups on the surface of the nanotubes and arrangement of their crystal lattice, the scanning electron microscopy (SEM; TSCAN Mira3 LMH) was used, and a Fourier transform

infrared (FT-IR) spectroscope (model: Tensor 27, Bruker, Germany) was utilized for the characterization of the surface properties and confirmation of the functional groups before and after the modification of the adsorbent within the range of 500-4,000  $\text{cm}^{-1}$ . In this research, X-ray diffraction XRD, dynamic light scattering (DLS), surface area (BET), and the zeta potential of the modified MWCNT nanocomposites were also determined.

Table 1. Properties of Direct Blue 15

Chemical structure	
Molecular formula	$\text{C}_{34}\text{H}_{24}\text{N}_6\text{Na}_4\text{O}_{16}\text{S}_4$
Type	Double azo class
Molecular weight	992.81 (g/mol)
$\lambda_{\text{max}}$	610 nm

### Adsorbent preparation

#### Modification of the MWCNTs with Ethylene Diamine Tetraacetic Acid (EDTA)

Initially, five grams of EDTA-2Na was dissolved in 100 milliliters of distilled water. Afterwards, 100 milliliters of the suspension containing one gram of the MWCNTs was slowly added at room temperature, and the solution was stirred on a high-speed magnetic stirrer for 3.5 hours.

#### Synthesis of the MWCNT composites modified with chitosan and EDTA

To synthesize the nanocomposite, two grams of chitosan was added to 100 milliliters of 3% acetic acid. Following that, the chitosan solution was mixed for one hour at high speed, and the solution was gently added to the prepared mixture of EDTA-2Na and MWCNTs at room temperature. To ensure complete dissolution, the solution was placed in a 37-kHz ultrasonic device for 20 min, followed by stirring at room temperature for 16 hours.

Finally, the solids were separated by centrifugation, washed repeatedly with distilled water, and dried under vacuum at the temperature of 60 °C.<sup>27</sup>

#### Adsorption experiments

Initially, the Direct Blue 15 dye powder was placed in oven at the temperature of 110 °C for two hours. Following that, the stock solution of the dye (1,000 mg/L) was prepared. Several parameters were assessed based on the pretests and results of similar studies, including pH, adsorbent dose, initial dye concentration, and contact time. The one-factor-at-a-time procedure was applied to optimize the parameters. To adjust the pH of the samples, 0.1-N NaOH and HCl were used. The removal efficiency of the system was recorded with the initial dye content of 10, 40, 60, 100, and 200 mg/L, pH of 3, 5, 7, 9, and 11, MWCNT dose of 0.25, 0.5, 1, 2, and 3 g/L and contact time of 10, 30, 45, 60, 90, and 120 min. In all the experiments, the stirrer speed was set at 200

rpm. After the centrifugation of the samples, the dye residue was measured using the UV/Vis spectrophotometer ( $\lambda_{\max}$ : 610 nm).<sup>28</sup> The dye removal efficiency was calculated using Eq. 1, as follows:

$$R(\%) = \frac{(c_0 - c_e)}{c_0} \times 100 \quad (1)$$

The adsorption capacity of the modified carbon nanotubes for various dye concentrations was determined using Eq. 2, as follows:

$$q_e = \frac{(c_0 - c_e)}{W} \quad (2)$$

where  $c_0$  and  $c_e$  are the initial and equilibrium concentrations (mg/L) of the dye after adsorption, respectively, and  $W$  shows the adsorbent mass in the solution volume (g/L).

## Results and Discussion

### Structure and morphology of the nanoparticles

Fig. 1 depicts the results of the FTIR of the modified MWCNTs. The FTIR spectroscopy of the virgin adsorbent (A), modified adsorbent (B), and virgin chitosan (C) has also been shown in the same figure, confirming the effective modification of the carbon nanotubes. In diagram B, the absorption peak was estimated at  $13,440 \text{ cm}^{-1}$  due to the presence of stretch O-H groups. The peaks observed within the range of  $2,922\text{-}2,904 \text{ cm}^{-1}$  could be suggestive of stretch C-H vibrations, which were amplified after the adsorbent modification. In diagram C, the peak of  $1,637 \text{ cm}^{-1}$  was observed in the presence of stretch  $\text{-C}=\text{C}$ -groups, and the peak corresponding to  $1,442 \text{ cm}^{-1}$  could express the presence of bend C-H groups and stretch C-Os. In addition, the peak observed at  $1,053 \text{ cm}^{-1}$  was related to crude chitosan, and the peaks corresponding to  $880 \text{ cm}^{-1}$  indicated the presence of C-H groups on the surface of the carbon nanotubes after modification.<sup>29, 30</sup>

Fig. 2 shows the SEM images of the CNTs and chitosan-modified MWCNT nanocomposites. Accordingly, the CNTs had low

porosity and tubular texture, while the composite surface was rough and porous. This characteristic of the composite surface is essential to surface adsorbents as it shows the expansion of the surface area and high absorption efficiency. In addition, the roughness and high porosity of the composite causes the hydrophobicity of the adsorbent surface.

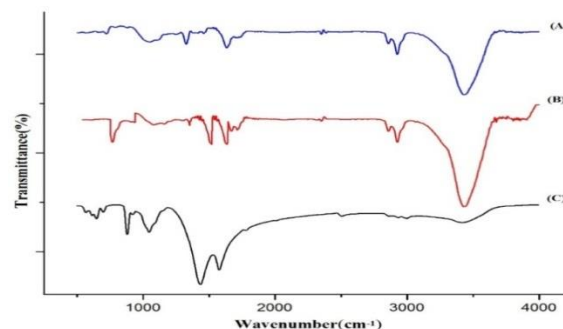


Fig. 1. FTIR spectroscopy of (A) virgin adsorbent, (B) modified adsorbent, and (C) virgin chitosan

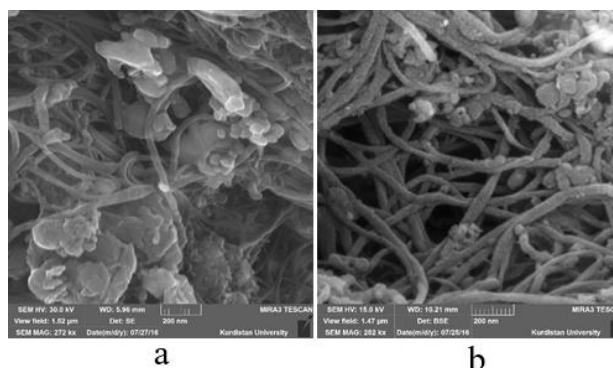


Fig. 2. SEM image of raw MWCNT and chitosan-modified MWCNT

Fig. 3 shows the XRD image of the CNTs and chitosan-modified MWCNT nanocomposites at the degree of  $10\text{-}90$ . The XRD images were employed to validate the crystalline structure and purity of the nanoparticles, measure the average distances between the layers and atomic series, and determine the location of the single crystal or grain. Although CNT is considered to be a non-crystalline material, its periodic structure leads to distinct X-ray scattering.<sup>31</sup> As depicted in the XRD scattering pattern, no significant peak was observed before the modification, with the

exception of the CNT peak. Through the modification and addition of chitosan, the peaks within the range of 30-50  $\text{cm}^{-1}$  were added. Figs. 4 and 5 depict the results of DLS and zeta potential for the CNTs and chitosan-modified MWCNT nanocomposites. The DLS revealed that the particle size in the raw CNTs was within the range of 120-300 nanometers, while it was approximately 160-500 nanometers in the prepared nanocomposites; as can be deduced, the addition of chitosan to the CNTs expanded the size.

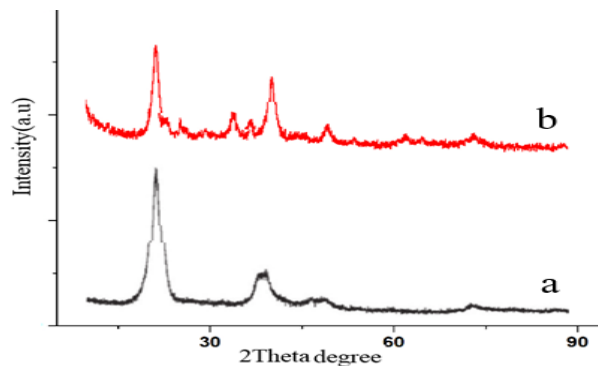


Fig. 3. XRD of (A) raw MWCNT and (B) chitosan-modified MWCNT nanocomposites

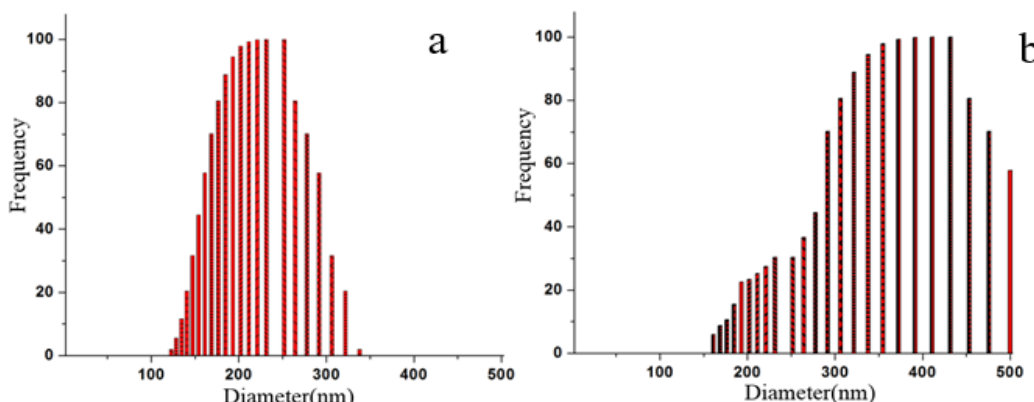


Fig. 4. DLS of (a) raw MWCNT and (b) chitosan-modified MWCNT nanocomposites

**Zeta potential**

The zeta potential is essential to the investigation of the properties of colloidal suspensions. In the present study, the electrical zeta potential of the raw MWCNT and chitosan-modified MWCNT nanocomposite was measured at the optimum pH. For this purpose, a suspension of 0.01 gram of the adsorbent in five milliliters of distilled water was prepared, and all the samples were irradiated with ultrasonic waves for 15 min before the zeta potential measurement. The results of the zeta potential (Fig. 5) confirmed that the zeta potential and motion increased in the contaminated nanoparticles. Therefore, the zeta potential in the raw MWCNT and chitosan-modified MWCNT nanocomposites was calculated to be 0.43 and -0.23 mV, respectively. Evidently, adsorption

modification enhanced both the surface area and surface charge.

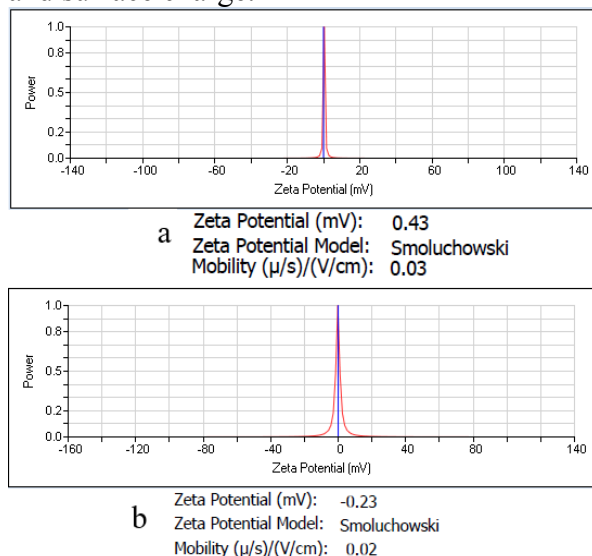


Fig. 5. Zeta potential of (A) raw MWCNT and (B) chitosan-modified MWCNT nanocomposites



### BET

Table 2 shows the results of the BET analysis for the raw MWCNT and chitosan-modified MWCNT nanocomposites. Accordingly, the composite holes and surface of the chitosan-modified MWCNT nanocomposite were more reliable compared to the raw MWCNTs, which in turn led to the reduction of the surface area. The findings of the current research in this regard are consisted with the previous studies.<sup>32</sup>

Table 2. BET of raw MWCNT and chitosan-modified MWCNT nanocomposites

Sample	SBET (m <sup>2</sup> /g)
Raw MWCNT	161
Chitosan-modified MWCNT nanocomposite	139

### Effects of various parameters on the adsorption process

#### pH solution

As is known, pH is a critical parameter in the adsorption process for the removal of environmental pollutants. Fig. 6 depicts the evaluation of the impact of pH on the removal efficiency of the Direct Blue 15 dye using the EDTA/chitosan-modified MWCNT adsorbent. With the increased pH, the removal efficiency decreased substantially, so that at pH=3, the dye removal efficiency was estimated at 99%, while it increased to 62% at pH=11. Therefore, it could be inferred that at acidic pH, the adsorbent had a higher efficiency in the removal of Direct Blue 15, while at the acidic pH=3, the adsorption efficiency was highest. The adsorption rate was observed to be high at neutral pH, and due to economic and environmental issues, pH=7 was selected as the optimum value.

In adsorption processes, the pH of the solution may play a key role in the adsorption capacity due to the impact of the solution chemistry, adsorbent surface charge, and functional groups. As is observed in Fig. 6, the

maximum acid removal occurred at acidic pH. Dye adsorption at high pH is due to the presence of hydroxyl functional groups and its association with the pH of the reaction medium. As is shown in Fig. 1 (B) and considering the presence of hydroxyl functional groups in the nanocomposite structure, the increased pH increased the bond between the OH groups on the adsorbent surfaces, resulting in the negative charge.<sup>33</sup> Since Direct Blue 15 is anionic with a negative charge in aqueous solutions and NaSO<sub>3</sub> group is also present in its structure, this dye is dissolved in water to the Na<sup>+</sup> cation and SO<sub>3</sub><sup>-</sup> anion, adsorbing the negative charge of the medium. Therefore, higher pH led to the increased negative charge adsorption locations and reduced positive charge sites, thereby decreasing the dye removal efficiency. Effective dye removal occurs when the positive surface charge environment is appropriate.<sup>34</sup> In a research conducted by Maleki *et al.* to eliminate reactive blue dye 19 using natural and modified zeolites, both sulfuric acid and phosphoric acid modified the sorbents, and the removal efficiency decreased with increased pH. This finding is consistent with the results of the present study.<sup>35</sup>

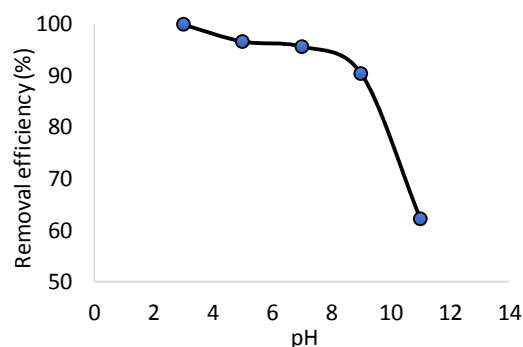


Fig. 6. Effect of solution pH on removal efficiency of Direct Blue 15 using EDTA/chitosan-modified MWCNT nanocomposite (adsorbent dose: 0.5 g/L, contact time: 60 min, initial dye concentration: 60 mg/L)

### Adsorbent dosage

In order to determine the impact of the adsorbent dosage on the dye removal, some experiments were performed with the adsorbent

dosage of 10-200 mg/L, pH of 7, contact time of 60 min, and initial dye content of 60 mg/L (Fig. 7). According to the obtained results, the increased adsorbent dose was associated with the higher efficiency, followed by a reduction. At the adsorbent dose of 0.5 g/L, the dye removal efficiency at the concentrations of 10 and 200 mg/L was estimated at 96 and 87%, respectively. However, increasing the adsorbent dosage to 3 g/L caused the removal efficiency to decrease by 91 and 77%, respectively. Consequently, the adsorbent dosage of 0.5 g/L was considered optimal. By increasing the adsorbent dosage to a certain level, the removal efficiency increased for all the dyes. When the adsorbent further increased, the dye removal efficiency remained constant and eventually declined. By increasing the adsorbent dosage from 0.25 to 0.5 g/L for the dye at the concentration of 60 mg/L, the removal efficiency increased from 76.62 to 90%, which could be due to the increment in the adsorbent dosage as it made more surface area and more adsorption sites available to the dye molecules, thereby increasing the dye removal efficiency.

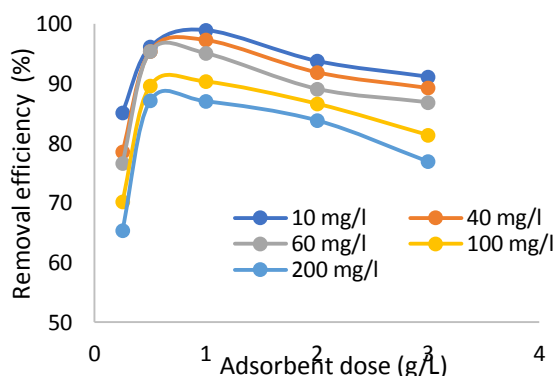


Fig. 7. Effect of adsorbent dose on removal efficiency using EDTA and chitosan-modified MWCNT (pH=7, contact time: 60 min, dye content: 60 mg/L)

Increasing the adsorbent dosage from 2 to 3 decreased the dye removal efficiency, which could be due to the fact that the gathering of adsorbent particles in high amounts led to interference and reduced surface area. In this regard, similar results have been reported by

Bazrafshan *et al.*<sup>36</sup> In another study by Ghaneian *et al.*, the efficacy of cuttlefish bone powder was investigated in the removal of the Reactive Blue 19 dye from aqueous solutions, and dye removal was reported to increase at higher adsorbent dosages due to the increased surface area and number of the adsorption sites.<sup>37</sup>

### Initial dye concentration

Fig. 8 shows the impact of the initial Direct Blue 15 concentration on the removal efficiency by the EDTA/chitosan-modified MWCNTs adsorbent. Fig. 2 depicts the correlation of dye adsorption ( $q_e$ ) with the increased concentration of the Direct Blue 15 dye. According to the findings, the dye removal efficiency declined with the increased initial dye concentration. However, the actual amount of the dye absorbed per unit mass of the nanotubes ( $q_e$ ) increased at higher initial dye concentrations. As can be seen in Fig. 5, at the initial dye concentration of 10 mg/L, the removal efficiency was approximately 96%, while at the concentration of 200 mg/L, the removal efficiency declined to approximately 87%. Furthermore, the removal efficiency decreased at higher initial dye concentrations. At the initial dye concentration of 10 mg/L, the removal efficiency was estimated at 96%, while at the concentration of 200 mg/L, the removal efficiency declined to approximately 87%. This reduction in the removal efficiency could be due to the saturation of the vacant active sites on the adsorbent surface at higher dye concentrations.<sup>38</sup>

At high dye concentrations, the quantity of the adsorbed dye on the carbon nanotubes increased (Diagram 2), so that the dye absorption capacity at the concentration of 10 mg/L was 19.22 mg/L, and at the concentration of 200 mg/L, it reached 348.32 mg/g; this could be attributed to the increment in the concentration gradient force.<sup>39</sup> In a study by Rahmani *et al.* regarding the removal of the

Reactive Black 5 dye using adsorption on efficiency decreased from 99 to 24% when the initial dye concentration increased from 50 to 300 mg/L.<sup>40</sup> This finding is consistent with the results of the present study. In another research by Luo *et al.* regarding the adsorption of natural red dye on halloysite nanotubes, the adsorption capacity increased from 24.92 to 65.69 mg/L with the increased dye concentration from 50 to 400 mg/L at 318 K. This is in line with the current research.<sup>41</sup>

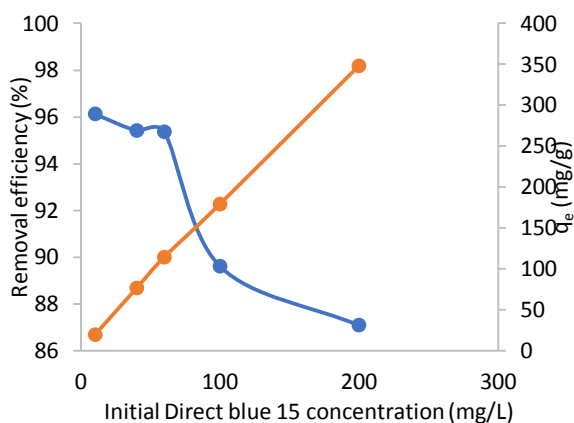


Fig. 8. Effect of initial dye concentration on removal efficiency using EDTA and chitosan-modified MWCNT (pH=7, contact time: 60 min, dye content: 60 mg/L)

### Contact time

Fig. 9 depicts the rate of dye absorption at different contact times. The obtained results confirmed that with the increased contact time, the dye removal efficiency initially increased, reaching balance within nearly one hour. During 10 min of contact, the dye adsorption was estimated at 104 mg/g, while with the increased contact time to 60 min, the adsorption rate reached 114 mg/g. Following that, the slope of the graph decreased, so that with the increased contact time to 120 min, the dye adsorption reached approximately 115 mg/g.

To assess the effect of contact time, the initial dye concentration of 60 mg/L, optimum pH (=7), and the obtained adsorbent dosage were considered. According to the findings, the dye adsorption capacity increased with the increased contact time, confirming that dye

anionic ion exchange resin, the removal adsorption occurred within less than one hour at high speed, so that the dye absorption rate of approximately 104 mg/g was observed at 10 min of exposure. This high amount was due to the fact that more unsaturated adsorption sites became available to the dye molecules during this contact duration, so that the dye absorption rate of approximately 104 mg/g was observed at 10 min of exposure. After the contact time of one hour, the slope of the graph decreased, so that with the increased contact time to 120 min, the dye adsorption reached approximately 115 mg/g.

The reduction of the adsorption rate with the increased contact time was due to the saturation of the active sites on the adsorbent surface and increased repulsive force. As a result, the equilibrium time for the adsorption of the Direct Blue 15 dye onto the MCNs was one hour in the present study. The findings reported by Shirmardi *et al.* are in line with the results of the present study.<sup>42</sup> Furthermore, the study by Malakootian *et al.* indicated that within the first 60 min, the uptake of anionic surfactant by modified Leica increased and reached balance within 90 min. However, the absorption rate did not vary significantly for more than 90 min,<sup>43</sup> which is consistent with the results of the present study.

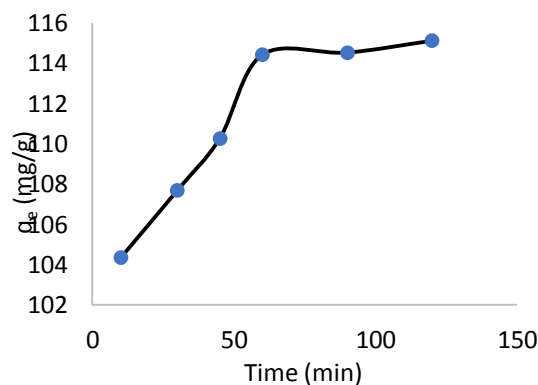


Fig. 9. Effect of contact time on removal efficiency using EDTA and chitosan-modified MWCNT (pH=7, adsorbent dose: 0.5 g/L, dye content 60 mg/L)

### Adsorption isotherms



Adsorption isotherms provide vital data to optimize the absorption system configuration. In the current research, the Langmuir and Freundlich models, which are among the most commonly employed adsorption isotherm models, were used to characterize the adsorption of the Direct Blue 15 dye onto the EDTA and chitosan-modified MWCNT.

The Langmuir isotherm is based on the premise of monolayer adsorption on a homogeneous structure adsorbent, assuming that all the sites have the same adsorption energy and are equal.<sup>44</sup> The linear form of the Langmuir equation is as Eq. 3:

$$\frac{C_e}{q_e} = \frac{1}{q_m k_L} + \frac{1}{q_m} C_e \quad (3)$$

where  $q_e$  is the amount of the adsorbed dye for a given amount of the consumed adsorbent (mg/g),  $C_e$  shows the equilibrium concentration of the solution (mg/L),  $q_m$  is the maximum amount of the dye required to form a singular

layer (mg/g), and  $k_L$  represents the Langmuir equilibrium absorbance constant (l/mg).

The Freundlich isotherm is an experimental equation associated with heterogeneous adsorbent surfaces, the linear form of which is expressed, as Eq. 4:

$$\ln(q_e) = \ln(K_f) + \frac{1}{n} \ln(C_e) \quad (4)$$

where the Freundlich  $k_f$  and  $n$  adsorption constants show the absorption capacity and intensity, respectively. In the current research, the results of the isotherm studies (Fig. 10, Table 3) revealed that the adsorption of the Direct Blue 15 dye on the EDTA and chitosan-modified MWCNT complied more with the Freundlich isotherm model, and its correlation-coefficient ( $R^2$ ) was higher than the Langmuir model. This finding indicated that the adsorption process was performed in the uncertain and heterogeneous locations of the adsorbent sites.

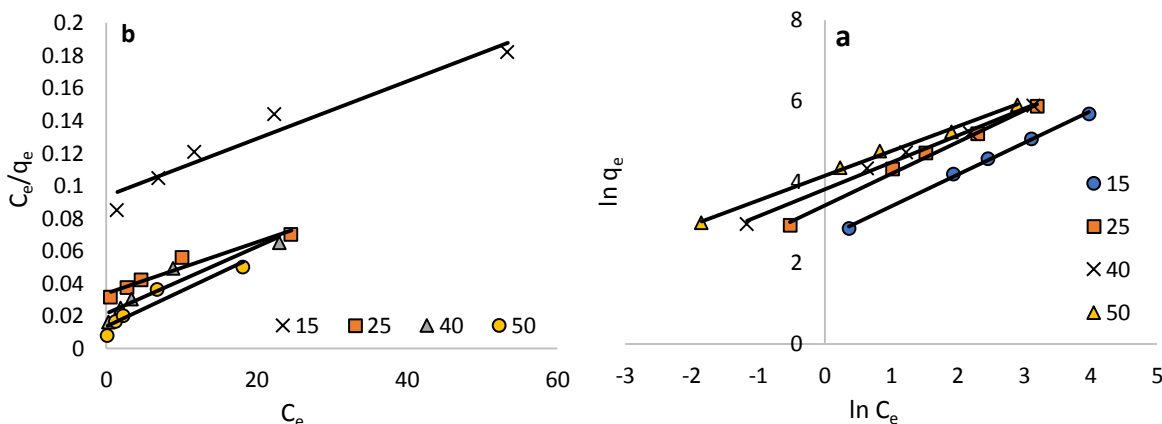


Fig. 10. Adsorption isotherms: (a) Freundlich model, (b) Langmuir model

Table 3. Variables of adsorption isotherms of Direct Blue 15 on EDTA and chitosan-modified MWCNT

T (°C)	Isotherm model						
	Frundlich			Langmuir			
	n	K <sub>f</sub>	R <sup>2</sup>	q <sub>m</sub>	R <sub>L</sub>	K <sub>L</sub>	R <sup>2</sup>
15	1.27	13.46	0.9975	555.55	0.46	0.019	0.9420
25	1.28	30.70	0.9945	625.00	0.26	0.047	0.9414
40	1.50	45.57	0.9940	500.00	0.15	0.092	0.9106
50	1.65	63.53	0.9957	454.54	0.09	0.160	0.9039

### Adsorption kinetics

In the current research, the equilibrium data were used to examine the reaction kinetics and their conformity with the pseudo-first- and pseudo-second-order kinetic equations. Eqs. 4 and 5 denote the first and second pseudo-linear Eqs., as follows:

$$\ln(q_e - q_t) = \ln(q_e) - k_1 t \quad (5)$$

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{1}{q_e} t \quad (6)$$

In Eqs. 5 and 6,  $k_1$  and  $k_2$  are the coefficients of velocity ( $\text{min}^{-1}$ ) and quadratic reaction constant ( $\text{g/mg min}$ ), respectively, and  $q_e$  and  $q_t$  show the adsorption capacity at the equilibrium and adsorption capacity at time  $t$ , respectively ( $\text{mg/g}$ ). The results and calculations of the kinetic equations in Table 3 and Figs. 9 and 10 indicate that the pseudo-second-order kinetic model was more accurate in describing the adsorption behavior of the dye molecules on the EDTA and chitosan-modified MWCNT. In a study by Nevine regarding the removal of the Direct Blue 106 dye using the activated carbon adsorbent made of pomegranate peel, the dye adsorption was reported to follow the pseudo-second-order kinetic model ( $R^2 \geq 0.99$ ).<sup>45</sup>

### Conclusion

According to the results, the modification of the MWCNT by chitosan and EDTA was significantly more effective in dye adsorption compared to the MWCNT. Therefore, MWCNT could be applied as a promising adsorbent for the removal of dyes from effluents.

### Acknowledgments

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